



# MBC – MRIDUL BHAIYA CLASSES

## TEST PAPER – 01

### PERIODIC CLASSIFICATION OF ELEMENTS

#### TEST

**MAXIMUM TIME: 35 MINUTES**

**MAXIMUM MARKS: 25 MARKS**

1. An element X belongs to group 1 and 3 period.
  - a. Give its electronic configuration
  - b. Formula when it combines with oxygen
  - c. State its nature (3)
2. An element to 3rd period and 17th group of the periodic table, find out
  - a. The electronic configuration of element on its left and right
  - b. Its valence electrons
  - c. Its size with in composition with neighboring elements (3)
3. State some characteristics of periodic table. (2)
4. Define atomic size? (1)
5. Name two isotopes of carbon. (1)
6. Name any two alkali present in group 1. (1)
7. An element X belongs to II group and 2nd period. Write the atomic number and name of element. (2)
8. The atomic number of X is 17, predict its
  - a. Valency
  - b. Formula of its halide
  - c. Type of ion formed
  - d. Reactivity with respect to the other of same group (4)
9. What happens to the valence electrons and valency as we move to left to right in a period and top to bottom in a period? (2)
10. Name the following elements
  - a. Two shells both of which are completely filled
  - b. Three shells with 2 valence electrons
  - c. Metals with valency 3, group number 13 and period 3. (3)
11. An element X from group 1 combines with an element Y from group 17 to form a compound. Both X and Y belongs to period II.
  - a. Give the formula of the compound formed
  - b. Will the compound be ionic or covalent
  - c. Give its electrons dot structure. (3)